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Fluorine-doped tin dioxide thin films prepared by chemical vapor deposition

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Transparent conductive tin dioxide thin films were prepared by a low-temperature atmospheric-pressure chemical vapor deposition method in air. The raw material was tin(II) trifluoroacetate. At a reaction temperature above 250 °C, polycrystalline thin films were obtained with a high deposition rate. This chemical vapor deposition method effectively incorporates F atoms into a crystalline structure and consequently maximizes the carrier concentration, yielding fluorine-doped SnO₂ films of very low resistivity. For the 1260-nm-thick film deposited at 400 °C, the resistivity was $5.92 \times 10^{-4} \Omega \text{ cm}$, and the sheet resistance was 4.69 Ω/\square . The deposition condition, structure, and characteristics of films were compared to the corresponding values in the chemical vapor depositions of SnO₂ from tin(II) acetate.

I. INTRODUCTION

Tin dioxide (SnO₂) is an *n*-type semiconductor which has many applications. In particular, SnO₂ thin films doped with antimony or fluorine are widely used in practice as transparent conductive films. The advantages of SnO₂ film are high chemical and mechanical stabilities even at high temperatures. These advantages find some applications in modern optoelectronic devices, such as solar cells.

High-quality SnO₂ film has been prepared by spray pyrolysis and chemical vapor deposition (CVD) methods.^{1,3} The CVD method yielded antimony-doped films^{1,2} with a resistivity of $8\text{--}10 \times 10^{-4} \Omega \text{ cm}$, and the fluorine-doped films³ with a resistivity of $4.5 \times 10^{-4} \Omega \text{ cm}$. The tin precursor, e.g., organotin^{1,2} and SnCl₄,³ and oxygen donor have been used with a dopant source, e.g., SbCl₃,¹ (CH₃)₃Sb,² and HF.³ In the CVD of doped SnO₂, therefore, the deposition processes have suffered from a lack of reproducibility because of the large number of parameters which have to be controlled.

In this paper, tin(II) bis-trifluoroacetate is proposed as a precursor for obtaining fluorine-doped SnO₂ (SnO₂:F) films. The low-temperature CVD of this precursor material, which involves bonds with fluorine atoms, yields fluorine-doped film with no need of supplying a fluorine donor. The preparation condition, structure, and characteristics of SnO₂:F film will be discussed by comparing those for SnO₂ films which were prepared from tin diacetate. Since the CVD of Sn(IV) dioxide from a Sn(II) compound includes oxidation processes, the SnO₂:F and SnO₂ films were also prepared by the CVD where O₂ gas was premixed with the source gas; and the effects of the O₂ gas premixing on deposition rate, crystallinity, and electric conductivity of film were discussed.

II. EXPERIMENT

Tin(II) trifluoroacetate [Sn(CF₃COO)₂, Nihon Kagaku Sangyo Co., Ltd.] was used as the source material for preparing SnO₂:F, and tin(II) acetate [Sn(CH₃COO)₂, Nihon Kagaku Sangyo Co., Ltd.] was used for preparing

SnO₂. Sn(CF₃COO)₂ and Sn(CH₃COO)₂ were heated at temperatures of 140–160 and 120–160 °C, respectively. The generated gases were entrained by carrier gases, which were mainly nitrogen. For O₂ premixing experiments in preparing SnO₂ film, N₂ gas for carrying Sn(CH₃COO)₂ was replaced by O₂ gas. Meanwhile, for O₂ premixing in preparing SnO₂:F film, O₂ gas was injected into pipeline (near the inlet of the reactor) where Sn(CF₃COO)₂ is carried with N₂ gas, because Sn(CF₃COO)₂ is reactive with O₂ even at low temperature.

A 76 × 26 mm borosilicate glass plate was used as the substrate, which was placed in the open-tube reactor heated by an external electric furnace. The reaction temperature ranged from 200 to 500 °C.

The composition of the film was measured by x-ray photoelectron spectroscopy. The crystallinity of the film was analyzed by the x-ray diffraction method with CuK α radiation. The electric resistivity and the Hall coefficient of the film were measured by the van der Pauw method. The optical transmittance of the film was obtained by means of a multipurpose recording spectrophotometer. A blank glass substrate was inserted into the reference beam path of the spectrophotometer.

III. RESULTS AND DISCUSSION

For two kinds of source materials, the deposition conditions were comparable; transparent conductive SnO₂ and SnO₂:F films were obtained in air at reaction temperatures above 200 and 250 °C, respectively. The films were not obtained in an inert (nitrogen) atmosphere. Thus, O₂ gas is inferred to make a role in oxidation of the decomposition product. The lower limit of the reaction temperature is much lower than that for organotin^{1,2} and SnCl₄.³

Figure 1 shows the Arrhenius plot of deposition rates of SnO₂, which were obtained at a source temperature of 120 °C and carrier gas flow rate of 1 ℓ/min . The deposition rates for two kinds of carrier gas agree fairly well. Thus, the O₂ premixing does not affect the deposition rate. The values of the deposition rate are of the same order of magnitude as those reported previously using the standard CVD methods. The

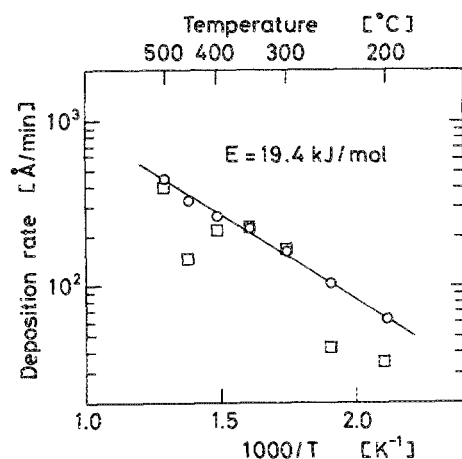


FIG. 1. Arrhenius plot of deposition rate of SnO_2 film prepared from $\text{Sn}(\text{CH}_3\text{COO})_2$ without premixing O_2 gas (○) and with premixing O_2 gas (□).

straight lines in this figure show that the activation energy is ~ 19.4 kJ/mol (0.20 eV/molecule), which is less than that (46.3–107.1 kJ/mol) obtained by the standard methods.² The preparation of $\text{SnO}_2\text{:F}$ film from tin(II) trifluoroacetate also showed the deposition rates of the same order of magnitude as that of SnO_2 films.

The x-ray diffraction patterns indicated that the crystallinity and the orientation of the films differed depending on kinds of both source material and carrier gas. Figures 2(a)–2(d) show typical examples of the x-ray diffraction pattern of the film on a borosilicate glass substrate at a reaction temperature of 350 °C. The $\text{SnO}_2\text{:F}$ film prepared from $\text{Sn}(\text{CF}_3\text{COO})_2$ without premixing O_2 gas shows textures in the [100] direction [Fig. 2(c)], but peak intensity of the [301] direction increases with increasing reaction temperature or increasing O_2 premixing [Fig. 2(d)]. In the meantime, SnO_2 film prepared from $\text{Sn}(\text{CH}_3\text{COO})_2$ without premixing O_2 gas showed a nonoriented polycrystalline nature [Fig. 2(a)] and the peak intensities of the (200) and (310) planes increased at 450 °C. Premixing O_2 gas into $\text{Sn}(\text{CH}_3\text{COO})_2$, however, yields a film of very weak crystallinity [Fig. 2(b)]. In short, for preparing SnO_2 from $\text{Sn}(\text{CH}_3\text{COO})_2$, premixed O_2 prevents crystalline growth, but for preparing $\text{SnO}_2\text{:F}$ from $\text{Sn}(\text{CF}_3\text{COO})_2$, premixed O_2 gas promotes (301) plane texturing, and F atoms in $\text{Sn}(\text{CF}_3\text{COO})_2$ promote (200) plane texturing. Thus, oxidation with premixed O_2 gas largely affects the crystalline growth, although it gives no effect on the growth rate.

Figure 3 shows the optical transmittances for the 536-nm-thick SnO_2 film and the 316-nm-thick $\text{SnO}_2\text{:F}$ film deposited on the borosilicate glass substrate. The transmittances are more than 85% (SnO_2 film) and 80% ($\text{SnO}_2\text{:F}$ film) in the visible range.

Figure 4 shows the electric resistivities of the SnO_2 and $\text{SnO}_2\text{:F}$ films as a function of reaction temperature, and Fig. 5 shows the resistivity of $\text{SnO}_2\text{:F}$ film as a function of the ratio of O_2 -to- N_2 gas flow rate. They were obtained with films prepared under the following conditions: source temperature = 140–160 °C, the total flow rate of gas = 0.4 ℓ /min, and the deposition time = 30–60 min. In Fig. 5, the

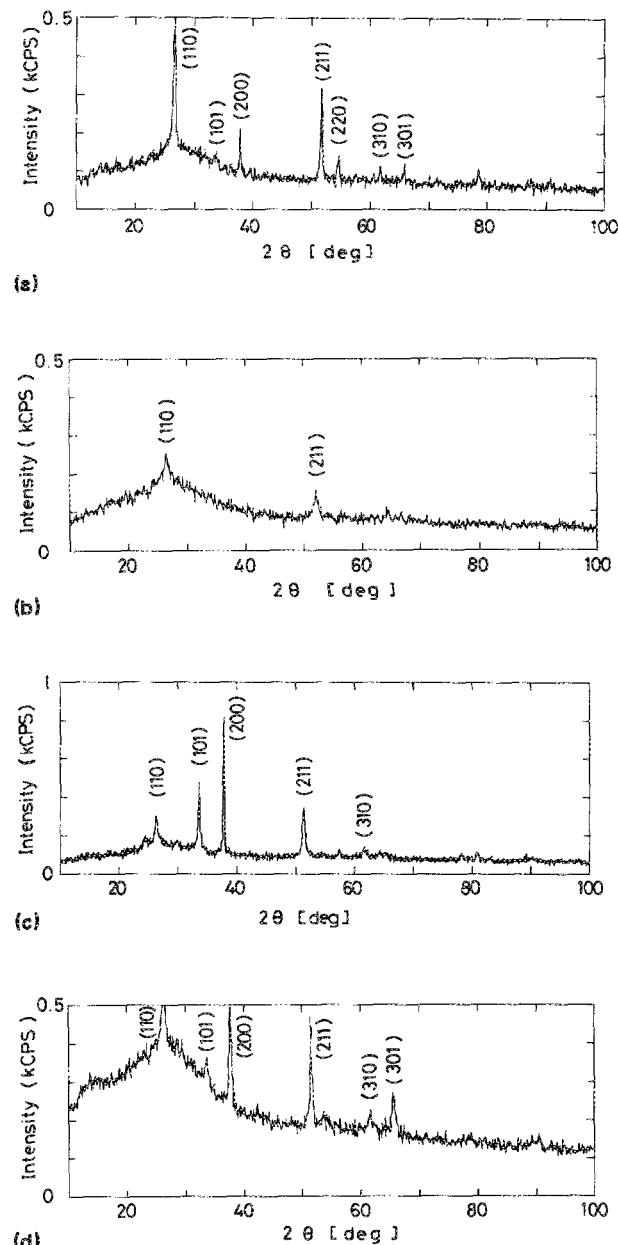


FIG. 2. X-ray diffraction pattern of film on borosilicate glass substrate: SnO_2 film prepared from $\text{Sn}(\text{CH}_3\text{COO})_2$ (a) without premixing O_2 gas and (b) with premixing O_2 gas; $\text{SnO}_2\text{:F}$ film prepared from $\text{Sn}(\text{CF}_3\text{COO})_2$ (c) without premixing O_2 gas and (d) with premixing O_2 gas.

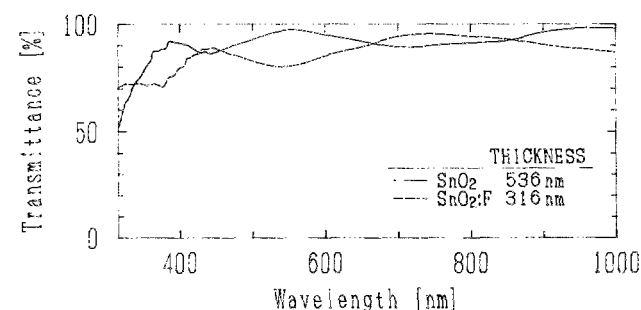


FIG. 3. Optical transmittances of SnO_2 film and $\text{SnO}_2\text{:F}$ film deposited on borosilicate glass substrate.

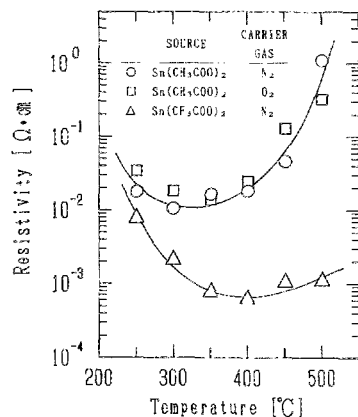


FIG. 4. Electric resistivities as a function of reaction temperature.

resistivity is found to be independent of the ratio O_2/N_2 under the wide range of $0 \leq O_2/N_2 \leq 1$. In Fig. 4 the results obtained with premixing O_2 gas agree well with the corresponding results without premixing O_2 gas. These facts indicate that the electric resistivity is not affected by the changes in crystalline structure caused by the O_2 premixing.

In Fig. 4, the resistivity of the $SnO_2:F$ film is much lower than that of the SnO_2 film, and the electric resistivity shows a strong dependence on reaction temperature. The lowest resistivity was obtained for the $SnO_2:F$ film deposited at 400 °C. After 60 min of deposition, the film thickness was 1260 nm. The resistivity was $5.92 \times 10^{-4} \Omega \text{ cm}$, and the sheet resistance was $4.69 \Omega/\square$.

The structure which gives high electric conductivity will be briefly discussed. Figures 6(a) and 6(b) show the carrier concentrations and Hall mobilities for both SnO_2 and $SnO_2:F$ films as a function of reaction temperature. It is noted that the carrier concentration for $SnO_2:F$ film is very high compared to values ($\sim 10^{20}$) reported previously^{1,3} for $SnO_2:F$ film. In the meantime, the F contents of the $SnO_2:F$ film were obtained by x-ray photoelectron spectroscopy. The values at reaction temperatures 350 and 450 °C were $0.730 \times 10^{21} \text{ cm}^{-3}$ (F/Sn = 0.0258) and $1.178 \times 10^{21} \text{ cm}^{-3}$ (F/Sn = 0.0278), respectively. Evidently, these values of F content are close to those of carrier concentration:

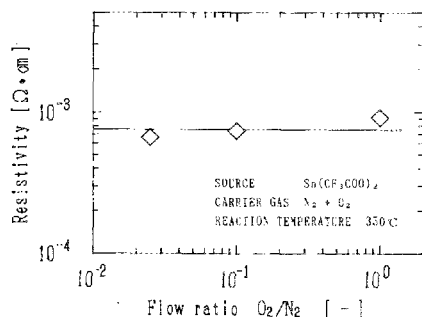
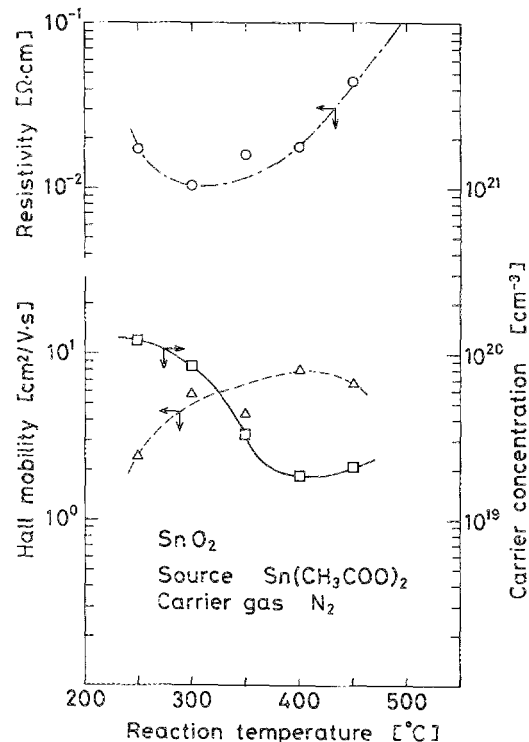
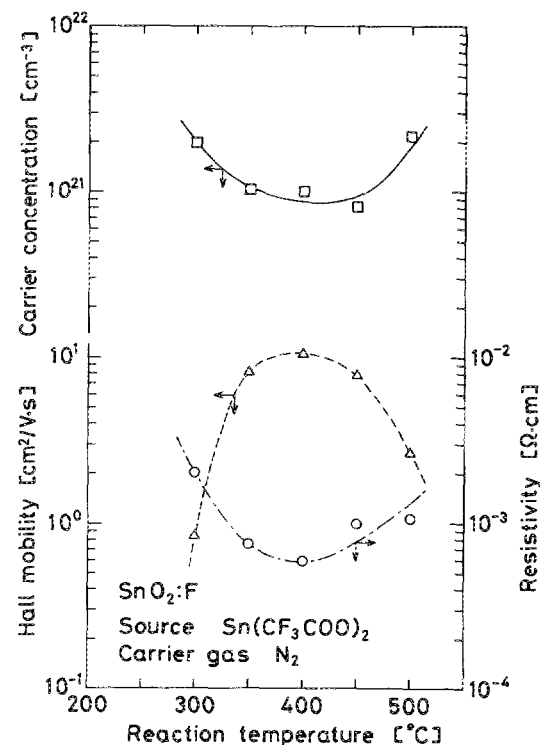


FIG. 5. Electric resistivity of $SnO_2:F$ film as a function of flow rate ratio of O_2 premixed in N_2 carrier.



(a)



(b)

FIG. 6. Electric resistivities, carrier concentration, and Hall mobilities as a function of reaction temperature. (a) SnO_2 film, (b) $SnO_2:F$ film.

$1.024 \times 10^{21} \text{ cm}^{-3}$ at 350 °C and $0.804 \times 10^{21} \text{ cm}^{-3}$ at 450 °C. Thus, CVD from tin bis-trifluoroacetate effectively incorporates F atoms into crystalline structure and consequently maximize the carrier concentration. On the other hand, Hall mobilities of $SnO_2:F$ film are not so high com-

pared to the reported value^{1,3} (~ 10), and they are of the same order of magnitude as that of SnO_2 film. These smaller values are inferred to be due to undevelopments in crystalline structure of the films which were formed under atmospheric pressure with rather high deposition rate. Thus, the resistivity was insensitive to the differences in crystallinity (for SnO_2) and texturing (for $\text{SnO}_2:\text{F}$) of the film in undeveloped state.

IV. CONCLUSIONS

Transparent conductive tin dioxide thin films were prepared by a low-temperature atmospheric-pressure chemical vapor deposition method in air. The raw material was tin(II) trifluoroacetate. At a reaction temperature above 250°C , polycrystalline thin films were obtained with a high deposition rate. This chemical vapor deposition method effectively incorporates F atoms into a crystalline structure and consequently maximizes the carrier concentration, yielding fluorine-doped SnO_2 films of very low resistivity.

For the 1260 nm-thick-film deposited at 400°C , the resistivity was $5.92 \times 10^{-4} \Omega \text{ cm}$, and the sheet resistance was $4.69 \Omega/\square$. The reaction temperature and deposition rate were comparable to the corresponding values in the chemical vapor depositions of SnO_2 from tin diacetate.

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